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This latest edition of CHEMISTRY: PRINCIPLES AND REACTIONS takes students directly to the crux of chemistry's fundamental concepts and allows you to efficiently cover all topics found in a typical general chemistry book. Based on the authors' extensive teaching experience, the book includes rigorous graded and concept-driven examples, as well as examples that focus on molecular

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reasoning and understanding. The Eighth Edition features a new and innovative example format, new talking labels within artwork, 25% new or revised problems, Chemistry: Beyond the Classroom essays that highlight some of the most up-to-date uses of chemistry, and end-of-chapter questions and Key Concepts that correlate to OWLv2, the #1 online homework and tutorial system for chemistry. Important Notice: Media content referenced within the product description or the product text may not be available in the ebook version. Fundamentals of Physical

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This book has been divided into four sections namely Physics, Chemistry, Biology and Computer Science, each divided into number of chapters as per the syllabi of General Science section asked in various competitive exams. The Physics section covers Motion, Force & Laws of Motion, Gravitation, Work, Energy & Power, Simple Harmonic Motion, Wave Motion, Light-Ray Optics, Current Electricity & Its Effects, Nuclear Physics, Semiconductor, Communication, etc whereas the Chemistry section has

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accompanying Interactive Examples in OWLv2 have been redesigned to better parallel the problem-solving approach in the narrative. New Capstone Problems have been added to a number of chapters. Important Notice: Media content referenced within the product description or the product text may not be available in the ebook version.

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mechanical operator for linear
momentum, angular momentum and
energy as Hermitian operator); The
average value of the square of Hermitian
operators; Commuting operators and
uncertainty principle(x & p ; E & t);
Schrodinger wave equation for a particle
in one dimensional box; Evaluation of
average position, average momentum and
determination of uncertainty in position
and momentum and hence Heisenberg's
uncertainty principle; Pictorial
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particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level; Lowest energy of the particle.

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Chapter 3. Chemical Dynamics – I: Effect of temperature on reaction rates; Rate law for opposing reactions of 1st order and 2nd order; Rate law for consecutive & parallel reactions of 1st order reactions; Collision

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theory of reaction rates and its limitations; Steric factor; Activated complex theory; Ionic reactions: single and double sphere models; Influence of solvent and ionic strength; The comparison of collision and activated complex theory. Chapter 4.

Electrochemistry – I: Ion-Ion

Interactions: The Debye-Huckel theory of ion-ion interactions; Potential and excess charge density as a function of distance from the central ion; Debye Huckel reciprocal length; Ionic cloud and its contribution to the total potential; Debye - Huckel limiting law of activity coefficients and its limitations; Ion-size effect on potential; Ion-size parameter and the theoretical mean-activity coefficient in the case of ionic clouds with finite-sized ions; Debye - Huckel-Onsager treatment for aqueous solutions and its limitations; Debye-Huckel-

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Onsager theory for non-aqueous solutions; The solvent effect on the mobility at infinite dilution; Equivalent conductivity (?) vs. concentration $c^{1/2}$ as a function of the solvent; Effect of ion association upon conductivity (Debye-Huckel - Bjerrum equation). Chapter 5. Quantum Mechanics – II: Schrodinger wave equation for a particle in a three dimensional box; The concept of degeneracy among energy levels for a particle in three dimensional box; Schrodinger wave equation for a linear harmonic oscillator & its solution by polynomial method; Zero point energy of a particle possessing harmonic motion and its consequence; Schrodinger wave equation for three dimensional Rigid rotator; Energy of rigid rotator; Space quantization; Schrodinger wave equation for hydrogen atom, separation of variable in polar spherical coordinates and its

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solution; Principle, azimuthal and magnetic quantum numbers and the magnitude of their values; Probability distribution function; Radial distribution function; Shape of atomic orbitals (s,p & d). Chapter 6. Thermodynamics – II: Clausius-Clayperon equation; Law of mass action and its thermodynamic derivation; Third law of thermodynamics (Nernst heat theorem, determination of absolute entropy, unattainability of absolute zero) and its limitation; Phase diagram for two completely miscible components systems; Eutectic systems, Calculation of eutectic point; Systems forming solid compounds $A_x B_y$ with congruent and incongruent melting points; Phase diagram and thermodynamic treatment of solid solutions. Chapter 7. Chemical Dynamics – II: Chain reactions: hydrogen-bromine reaction, pyrolysis of acetaldehyde,

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decomposition of ethane; Photochemical reactions (hydrogen - bromine & hydrogen -chlorine reactions); General treatment of chain reactions (ortho-para hydrogen conversion and hydrogen - bromine reactions); Apparent activation energy of chain reactions, Chain length; Rice-Herzfeld mechanism of organic molecules decomposition(acetaldehyde); Branching chain reactions and explosions (H₂-O₂ reaction); Kinetics of (one intermediate) enzymatic reaction : Michaelis-Menton treatment; Evaluation of Michaelis 's constant for enzyme-substrate binding by Lineweaver-Burk plot and Eadie-Hofstae methods; Competitive and non-competitive inhibition. Chapter 8. Electrochemistry – II: Ion Transport in Solutions: Ionic movement under the influence of an electric field; Mobility of ions; Ionic drift velocity and its relation with current

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density; Einstein relation between the absolute mobility and diffusion coefficient; The Stokes- Einstein relation; The Nernst -Einstein equation; Walden's rule; The Rate-process approach to ionic migration; The Rate process equation for equivalent conductivity; Total driving force for ionic transport, Nernst - Planck Flux equation; Ionic drift and diffusion potential; the Onsager phenomenological equations; The basic equation for the diffusion; Planck-Henderson equation for the diffusion potential.

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1. This book is based on CBSE's new syllabus and directives (2022-2023). All of the basic concepts & NCERT Textbook's answers are included. 2. Additionally, it includes previous year board questions, Competency-based questions, and NCERT Exemplars. 3. For a full revision of the curriculum, all types of questions are offered, including

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Multiple Choice Questions, Assertion-Reason Questions, Case-based Questions, Source-based questions, Passage-based Questions, Very Short Answer Questions, Short Answer Questions, and Long Answer Questions. 4. Solved CBSE Sample Papers and Exam Papers for Terms 1 and 2 (2021-22) are included to assist students in their Exam Preparation Fundamentals of Physical Chemistry