

Butler Ionic Equilibrium Solubility Chapter 6

Chemistry is covered at just the right depth for students to develop a thorough understanding of natural processes.Chemical processes shape the world we live in; the air we breathe, the water we drink, the weather we experience. Guiding us through the chemical composition of the three key environmental systems; the atmosphere, hydrosphere, and terrestrial environment; the authors explain the chemical processeswhich occur within and between each system, allowing for better understanding of how they behave. We then see how human activity continues to affect the chemical behaviour of these environmental systems, and what the consequences of these natural processes being disturbed can be.

This book provides a modern and easy-to-understand introduction to the chemical equilibria in solutions. It focuses on aqueous solutions, but also addresses non-aqueous solutions, covering acid–base, complex, precipitation and redox equilibria. The theory behind these and the resulting knowledge for experimental work build the foundations of analytical chemistry. They are also of essential importance for all solution reactions in environmental chemistry, biochemistry and geochemistry as well as pharmaceuticals and medicine. Each chapter and section highlights the main aspects, providing examples in separate boxes. Questions and answers are included to facilitate understanding, while the numerous literature references allow students to easily expand their studies.

Martin's Physical Pharmacy and Pharmaceutical Sciences is considered the most comprehensive text available on the application of the physical, chemical and biological principles in the pharmaceutical sciences. It helps students, teachers, researchers, and industrial pharmaceutical scientists use elements of biology, physics, and chemistry in their work and study. Since the first edition was published in 1960, the text has been and continues to be a required text for the core courses of Pharmaceutics, Drug Delivery, and Physical Pharmacy. The Sixth Edition features expanded content on drug delivery, solid oral dosage forms, pharmaceutical polymers and pharmaceutical biotechnology, and updated sections to cover advances in nanotechnology.

Solubility is fundamental to most areas of chemistry and is one of the most basic of thermodynamic properties. It underlies most industrial processes. Bringing together the latest developments and ideas, Developments and Applications in Solubility covers many varied and disparate topics. The book is a collection of work from leading experts in their fields and covers the theory of solubility, modelling and simulation, industrial applications and new data and recent developments relating to solubility. Of particular interest are sections on: experimental, calculated and predicted solubilities; solubility phenomena in 'green' quaternary mixtures involving ionic liquids; molecular simulation approaches to solubility; solubility impurities in cryogenic liquids and carbon dioxide in chemical processes. The book is a definitive and comprehensive reference to what is new in solubility and is ideal for researcher scientists, industrialists and academics

The Best Books for Academic Libraries: Science, technology, and agriculture

Physicochemical Principles

Physical Chemical and Biopharmaceutical Principles in the Pharmaceutical Sciences

A Mathematical Approach

Ionic Equilibrium

Chemical Principles

Discussing a comprehensive range of topics, Advanced Pharmaceutics: Physicochemical Principles reviews all aspects of physical pharmacy. The book explains the basic, mechanistic, and quantitative interpretation skills needed to solve physical pharmacy related problems. The author supplies a strong fundamental background and extensively covers them

Acids and bases are ubiquitous in chemistry. Our understanding of them, however, is dominated by their behaviour in water. Transfer to non-aqueous solvents leads to profound changes in acid-base strengths and to the rates and equilibria of many processes: for example, synthetic reactions involving acids, bases and nucleophiles; isolation of pharmaceutical actives through salt formation; formation of zwitter- ions in amino acids; and chromatographic separation of substrates. This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non-aqueous solvents. The behaviour is related where possible to that in water, but correlations and contrasts between solvents are also presented. Fundamental background material is provided in the initial chapters: quantitative aspects of acid-base equilibria, including definitions and relationships between solution pH and species distribution; the influence of molecular structure on acid strengths; and acidity in aqueous solution. Solvent properties are reviewed, along with the magnitude of the interaction energies of solvent molecules with (especially) ions; the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial. Experimental methods for determining dissociation constants are described in detail. In the remaining chapters, dissociation constants of a wide range of acids in three distinct classes of solvents are discussed: protic solvents, such as alcohols, which are strong hydrogen-bond donors; basic, polar aprotic solvents, such as dimethylformamide; and low-basicity and low polarity solvents, such as acetonitrile and tetrahydrofuran. Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents, and there is a strong differentiation between the response of neutral and charged acids to solvent change. Ion-pairing and hydrogen-bonding equilibria, such as between phenol and phenoxide ions, play an increasingly important role as the solvent polarity decreases, and their influence on acid-base equilibria and salt formation is described.

*** The present work is designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry, biochemistry, geochemistry, and interdisciplin ary environmental fields. The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature. The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems, instead of the traditional approach through approximate concentrations and idealized, infinite-dilution assumptions. There is repeated emphasis on the use of corrected, conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and/or against graphs of species distributions. Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations. The coverage of solution equilibrium phenomena is not, however, exhaustively comprehensive. Rather, I have chosen to offer funda mental and rigorous examinations of homogeneous step-equilibria and their interactions with solubility and redox equilibria. Many examples are worked out in detail to demonstrate the use of equi librium calculations and diagrams in various fields of investigation.**

The aim of this book is to provide an overview of the importance of stoichiometry in the biomedical field. It proposes a collection of selected research articles and reviews which provide up-to-date information related to stoichiometry at various levels. The first section deals with host-guest chemistry, focusing on selected calixarenes, cyclodextrins and crown ethers derivatives. In the second and third sections the book presents some issues concerning stoichiometry of metal complexes and lipids and polymers architecture. The fourth section aims to clarify the role of stoichiometry in the determination of protein interactions, while in the fifth section some selected experimental techniques applied to specific systems are introduced. The last section of the book is an attempt at showing some interesting connections between biomedicine and the environment, introducing the concept of biological stoichiometry. On this basis, the present volume would definitely be an ideal source of scientific information to researchers and scientists involved in biomedicine, biochemistry and other areas involving stoichiometry evaluation.

Principles of Modern Chemistry

Fundamentals of Analytical Chemistry

Groundwater Science

Environmental Chemistry, Eighth Edition

Solubility and pH Calculations

Electrochemical Methods: Fundamentals and Applications, 2nd Edition

Advanced Pharmaceutics

Solvent Effects on Acid-Base Strength

Surfactants in Precision Cleaning

Journal

th The technology of froth flotation, invented in the early 20 century was first used for the concentration of sulfide minerals. Since then it has been applied for the processing of many nonsulfide ores as weIl, including oxides, carbonates, silicates, soluble minerals like halite and sylvite and energy minerals like coal and bitumen. In recent years it has been used for several nonrnineral applications, such as waste water treatment, deinking of paper for recycling and resource recovery from industrial wastes he technology continues to grow with new applications reported every year. Flotation is based on chemical phenomena occurring at the interfaces, solid/water and air/water. Surface Chemistry principles have played a significant role in the development of flotation technology. Knowledge of aqueous solution chemistry and electrochemistry has added to our understanding of the reactions in flotation systems. Professor Jan Leja's book has well served researchers and students as they tried to understand the chemistry of flotation, and it is a significant contribution to the advancement of knowledge. However, since the book was first published, new research techniques and ever growing information have made an update necessary. The revised edition compiled by Dr. S. R. Rao has brought together fundamental aspects of the chemistry of flotation and how they apply to practical systems. It should serve all who are working in the area of flotation and interested in exploring new applications of flotation technology.

Environmental Chemistry, Eighth Edition builds on the same organizational structure validated in previous editions tosystematically develop the principles, tools, and techniques of environmental chemistry to provide students and professionals with a clear understanding of the science and its applications. Revised and updated since the publication of the best-selling Seventh Edition, this text continues to emphasize the major concepts essential to the practice of environmental science, technology, and chemistry while introducing the newest innovations to the field. The author provides clear explanations to important concepts such as the anthosphere, industrial ecosystems, geochemistry, aquatic chemistry, and atmospheric chemistry, including the study of ozone-depleting chlorofluorocarbons. The subject of industrial chemistry and energy resources is supported by pertinent topics in recycling and hazardous waste. Several chapters review environmental biochemistry and toxicology, and the final chapters describe analytical methods for measuring chemical and biological waste. New features in this edition include: enhanced coverage of chemical fate and transport; industrial ecology, particularly how it is integrated with green chemistry; conservation principles and recent accomplishments in sustainable chemical science and technology; a new chapter addressing terrorism and threats to the environment; and the use of real world examples.

Transport and transformation processes are key for determining how humans and other organisms are exposed to chemicals. These processes are largely controlled by the chemicals’ physical-chemical properties. This new edition of the Handbook of Physical-Chemical Properties and Environmental Fate for Organic Chemicals is a comprehensive series in four volumes that serves as a reference source for environmentally relevant physical-chemical property data of numerous groups of chemical substances. The handbook contains physical-chemical property data from peer-reviewed journals and other valuable sources on over 1200 chemicals of environmental concern. The handbook contains new data on the temperature dependence of selected physical-chemical properties, which allows scientists and engineers to perform better chemical assessments for climatic conditions outside the 20–25-degree range for which property values are generally reported. This second edition of the Handbook of Physical-Chemical Properties and Environmental Fate for Organic Chemicals is an essential reference for university libraries, regulatory agencies, consultants, and industry professionals, particularly those concerned with chemical synthesis, emissions, fate, persistence, long-range transport, bioaccumulation, exposure, and biological effects of chemicals in the environment. This resource is also available on CD-ROM

Undergraduate-level text focuses on three lines of the development of contemporary chemical structural theory: the classical theory of bonding in molecules; the ionic interpretation of electrolyte solutions; and the physical theory of atomic structure. 186 illustrations. 1969 edition.

UNESCO-IHE PhD Thesis

The Theory of Acid-Base, Complex, Precipitation and Redox Equilibria

Geological Survey Water-supply Paper

Chemical Kinetics and Reaction Dynamics

Removal of Contaminants at the Micro and Nanoscale

Inputs, Fates, and Effects

The Medicinal Chemist’s Guide to Solving ADMET Challenges summarizes a series of design strategies and tactics that have been successfully employed across pharmaceutical and academic laboratories to solve common ADMET issues. These are exemplified with a curated collection of concrete examples displayed in a highly visual “table-of-contents” style format, allowing readers to rapidly identify the most promising approaches applicable to their own challenges. Each ADMET parameter is introduced in a concise yet comprehensive manner and includes background, relevance and screening strategies. Medicinal chemistry knowledge of how best to modify molecular structure to solve ADMET issues is challenging to retrieve from the literature, public databases and even corporate data warehouses. The Medicinal Chemist’s Guide to Solving ADMET Challenges addresses this gap by presenting state-of-the-art design strategies put together by a global group of experienced medicinal chemists and ADMET experts across academia and the pharmaceutical industry.

Groundwater Science is a timely, current, and comprehensive presentation of groundwater hydrology that integrates chemistry, physics, geology, and calculus. With the input of students and other hydrology instructors, the author has developed a text reference that will be appreciated by students and professors alike. The book covers recent contamination topics such as non-aqueous-phase liquids (NAPLs), complex solute transport processes, and remediation. It contains clear explanations, two-color figures, case studies, and worked examples throughout. This book is recommended for upper-level undergraduate and graduate students in hydrology, geology, environmental science, and civil engineering departments as well as for scientists and engineers in the groundwater field. Integrates up-to-date material on field methods and flow modeling methods Covers recent contamination topics including non-aqueous-phase liquids (NAPLs), complex solute transport processes, and remediation Presents clear explanations, two-color figures, case studies, and worked examples throughout

Pharmaceutics is one of the most diverse subject areas in all of pharmaceutical science. In brief, it is concerned with the scientific and technological aspects of the design and manufacture of dosage forms or medicines. An understanding of pharmaceutics is therefore vital for all pharmacists and those pharmaceutical scientists who are involved with converting a drug or a potential drug into a medicine that can be delivered safely, effectively and conveniently to the patient. Now in its fourth edition, this best-selling textbook in pharmaceutics has been brought completely up to date to reflect the rapid advances in delivery methodologies by eye and injection, advances in drug formulations and delivery methods for special groups (such as children and the elderly), nanomedicine, and pharmacognosy. At the same time the editors have striven to maintain the accessibility of the text for students of pharmacy, preserving the balance between being a suitably pitched introductory text and a clear reflection of the state of the art. provides a logical, comprehensive account of drug design and manufacture includes the science of formulation and drug delivery designed and written for newcomers to the design of dosage forms New to this edition New editor: Kevin Taylor, Professor of Clinical Pharmaceutics, School of Pharmacy, University of London. Twenty-two new contributors. Six new chapters covering parenteral and ocular delivery; design and administration of medicines for the children and elderly; the latest in plant medicines; nanotechnology and nanomedicines, and the delivery of biopharmaceuticals. Thoroughly revised and updated throughout.

Long considered the standard for honors and high-level mainstream general chemistry courses, PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern, rigorous, and chemically and mathematically accurate text on the market. This authoritative text features an "atoms first" approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure (Chapter 6), Electrochemistry (Chapter 17), and Molecular Spectroscopy and Photochemistry (Chapter 20). In addition, the text utilizes mathematically accurate and artistic atomic and molecular orbital art, and is student friendly without compromising its rigor. End-of-chapter study aids focus on only the most important key objectives, equations and concepts, making it easier for students to locate chapter content, while applications to a wide range of disciplines, such as biology, chemical engineering, biochemistry, and medicine deepen students' understanding of the relevance of chemistry beyond the classroom.

Quantitative Chemical Analysis

Chemical Equilibria in Analytical Chemistry

A Brief Introduction to Quantitative Chemical Analysis

Acids and Bases

Oil in the Sea III

Faraday transactions. I.

Surfactants in Precision Cleaning: Removal of Contaminants at the Micro and Nanoscale is a single source of information on surfactants, emulsions, microemulsions and detergents for removal of surface contaminants at the micro and nanoscale. The topics covered include cleaning mechanisms, effect of surfactants, types of stable dispersions (emulsions, microemulsions, surfactants, detergents, etc.), cleaning technology, and cleaning applications. Users will find this volume an excellent resource on the use of stable dispersions in precision cleaning. Single source of current information on surfactants, emulsions, microemulsions and detergents for precision cleaning applications Includes a list of extensive reference sources Discusses specific selection and properties of surfactants and their use in cleaning Provides a guide for cleaning applications in different industry sectors

Discover the principles and practices behind analytic chemistry as you study its applications in medicine, industry and the sciences with Skoog/West/Holler/Crouch's FUNDAMENTALS OF ANALYTICAL CHEMISTRY, 10th Edition. This award-winning author team presents the latest developments in analytic chemistry today using a reader-friendly yet systematic and thorough approach. Each chapter begins with a compelling story and stunning visuals. Dynamic photos from renowned chemistry photographer Charlie Winters capture attention while reinforcing key principles. New features highlight chemistry-related careers. You also learn how to use Excel 2019 as a problem-solving tool in analytical chemistry with new exercises, updates and examples. Important Notice: Media content referenced within the product description or the product text may not be available in the ebook version.

A comprehensive and detailed study on the scaling potential of calcium carbonate in seawater reverse osmosis systems (SWRO), this book provides a new approach for calculating the degree of supersaturation and the pH of the SWRO systems concentrates with the assistance of the feed-water pH and the inorganic carbon constituents. Furthermore, the book highlights the weakness in the present supersaturation indices and membrane manufacturers programs. Finally, the research suggested that SWRO concentrate is much lower undersaturated with respect to calcium carbonate than previously thought. This was confirmed by comprehensive pilot testing where acids and antiscalants used to prevent calcium carbonate scaling were completely eliminated from the pilot plant.

Solvent systems are integral to drug development and pharmaceutical technology. This single topic encompasses numerous allied subjects running the gamut from recrystallization solvents to biorelevant media. The goal of this contribution to the AAPS Biotechnology: Pharmaceutical Aspects series is to generate both a practical handbook as well as a reference allowing the reader to make effective decisions concerning the use of solvents and solvent systems. To this end, the monograph was created by inviting recognized experts from a number of fields to author relevant sections. Specifically, 15 chapters have been designed covering the theoretical background of solubility, the effect of ionic equilibria and pH on solubilization, the use of solvents to effect drug substance crystallization and polymorph selection, the use of solvent systems in high throughput screening and early discovery, solvent use in preformulation, the use of solvents in bio-relevant dissolution and permeation experiments, solvents and their use as toxicology vehicles, solubilizing media and excipients in oral and parenteral formulation development, specialized vehicles for protein formulation and solvent systems for topical and pulmonary drug administration. The chapters are organized such that useful decision trees are included together with the scientific underpinning for their application. In addition, trends in the use of solvent systems and a balance of current views make this monograph useful to both the novice and experienced researcher and to scientists at all developmental stages from early discovery to late pharmaceutical operations.

The Development of Chemical Principles

A Global Perspective

The Principles of Chemical Equilibrium

Volume 1: Fundamentals

Carbon Dioxide Equilibria and Their Applications

Surface Chemistry of Froth Flotation

Measurement of equilibrium state; Measurement of equilibrium constants; Mathematical methods used in equilibrium calculations; Strong acids and bases; Weak monoprotic acids and bases; Precipitation and the Solubility product; Polyprotic acids; Introduction to complex formation equilibria; Complex formation; advanced topics; Oxidation-reduction equilibria; Nonideality corrections.

This book of general analytical chemistry – as opposed to instrumental analysis or separation methods – in aqueous solutions is focuses on fundamentals, which is an area too often overlooked in the literature. Explanations abound of the chemical and physical principles of different operations of chemical analysis in aqueous solutions. Once these principle are firmly established, numerous examples of applications are also given.

Carbon dioxide is the most important greenhouse gas after water vapor in the atmosphere of the earth. More than 98% of the carbon of the atmosphere-ocean system is stored in the oceans as dissolved inorganic carbon. The key for understanding critical processes of the marine carbon cycle is a sound knowledge of the seawater carbonate chemistry, including equilibrium and nonequilibrium properties as well as stable isotope fractionation. Presenting the first coherent text describing equilibrium and nonequilibrium properties and stable isotope fractionation among the elements of the carbonate system. This volume presents an overview and a synthesis of these subjects which should be useful for graduate students and researchers in various fields such as biogeochemistry, chemical oceanography, paleoceanography, marine biology, marine chemistry, marine geology, and others. The volume includes an introduction to the equilibrium properties of the carbonate system in which basic concepts such as equilibrium constants, alkalinity, pH scales, and buffering are discussed. It also deals with the nonequilibrium properties of the seawater carbonate chemistry. Whereas principle of chemical kinetics are recapitulated, reaction rates and relaxation times of the carbonate system are considered in details. The book also provides a general introduction to stable isotope fractionation and describes the partitioning of carbon, oxygen, and boron isotopes between the species of the carbonate system. The appendix contains formulas for the equilibrium constants of the carbonate system, mathematical expressions to calculate carbonate system parameters, answers to exercises and more.

Chemical Kinetics and Reaction Dynamics brings together the major facts and theories relating to the rates with which chemical reactions occur from both the macroscopic and microscopic point of view. This book helps the reader achieve a thorough understanding of the principles of chemical kinetics and includes: Detailed stereochemical discussions of reaction steps Classical theory based calculations of state-to-state rate constants A collection of matters on kinetics of various special reactions such as micellar catalysis, phase transfer catalysis, inhibition processes, oscillatory reactions, solid-state reactions, and polymerization reactions at a single source. The growth of the chemical industry greatly depends on the application of chemical kinetics, catalysts and catalytic processes. This volume is therefore an invaluable resource for all academics, industrial researchers and students interested in kinetics, molecular reaction dynamics, and the mechanisms of chemical reactions.

Ionic Equilibria in Analytical Chemistry

Solvent Systems and Their Selection in Pharmaceutics and Biopharmaceutics

Essentials of General, Organic and Biochemistry

The Design and Manufacture of Medicines

Handbook of Physical-Chemical Properties and Environmental Fate for Organic Chemicals, Second Edition

With Applications in Chemistry and Chemical Engineering

Ionic EquilibriumSolubility and pH CalculationsJohn Wiley & Sons

A celebrated classic in the field updated and expanded to includethe latest computerized calculation techniques In 1964, James N. Butler published a book in which he presentedsome simple graphical methods of performing acid-base, solubility,and complex formation equilibrium calculations. Today, both thebook and these methods have become standard for generations ofstudents and professionals in fields ranging from environmentalscience to analytical chemistry. Named a "Citation Classic" by theScience Citation Index in 1990, the book, Ionic Equilibrium,continues to be one of the most widely used texts on the subject.So why tamper with near-perfection by attempting a revision of thatclassic? The reason is simple-- the recent rapid development andwide availability of personal computers. In the revised Ionic Equilibrium, Dr. Butler updates his 1964 workby abandoning the slide rule and graph paper for the PCspreadsheet. He also expands the original coverage with extensivematerial on basic principles and recent research. The first part of Ionic Equilibrium is devoted to the fundamentalsof acid-base, solubility, and complex formation equilibria. In thesecond part, the author discussees oxidation-reduction equilibria,develops the principles of carbon dioxide equilibria, presents casestudies demonstrating the ways in which carbon dioxide equilibriaare used in physiology and oceanography, and explores thepossibility of a pH scale for brines. The concluding chapter, written by David R. Cogley, gives examples of general computerprograms that are capable of performing equilibrium calculations onsystems of many components. Replete with real-world examples, details of importantcalculations, and practical problems, Ionic Equilibrium is an idealcourse text for students of environmental chemistry, engineering,or health; analytical chemistry; oceanography; geochemistry;biochemistry; physical chemistry; and clinical chemistry. It isalso a valuable working resource for professionals in those fieldsas well as industrial chemists involved with solution chemistry.

A broad and comprehensive survey of the fundamentals for electrochemical methods now in widespread use. This book is meant as a textbook, and can also be used for self-study as well as for courses at the senior undergraduate and beginning graduate levels. Knowledge of physical chemistry is assumed, but the discussions start at an elementary level and develop upward. This revision comes twenty years after publication of the first edition, and provides valuable new and updated coverage.

Sample Text

Environmental Chemistry

CO2 In Seawater: Equilibrium, Kinetics, Isotopes

Martin's Physical Pharmacy and Pharmaceutical Sciences

Stoichiometry and Research

Developments and Applications in Solubility

A Practical Introduction for the Physical and Life Sciences